The *in vitro* effect of zinc on the inhibition of human δ-aminolevulinic acid dehydratase by lead

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Border, E. A., Cantrell, A. C., and Kilroe-Smith, T. A. (1976). *British Journal of Industrial Medicine*, 33, 85-87. The *in vitro* effect of zinc on the inhibition of human δ-aminolevulinic acid dehydratase by lead. The antagonistic effect of zinc on inhibition of human aminolevulinic acid dehydratase (ALAD) by lead was examined *in vitro*. The phenomenon was studied at pH 6.5-7.5. Zinc and lead were added at concentrations ranging from physiological levels to levels far in excess of those expected in heavy industrial exposure. ALAD activity of normal blood assayed in the presence of zinc was unaffected by exogenous lead if the added zinc concentration was above 0.53 mmol/l blood. If zinc was added at 0.15 mmol/l blood, the total zinc concentration was approximately 0.28 mmol/l blood, or double the normal endogenous level, and was of the order found in heavy industrial exposure to zinc. In this case, ALAD assay values were appreciably raised and might have fallen into the normal range in spite of a dangerous total blood lead of over 0.0048 mmol/l. When zinc was added *in vitro* to blood from a worker with a blood lead level of 0.0043 mmol/l the ALAD values obtained were also raised and did not reflect the dangerous level of lead in the patient.

Recent reports (Abdulla and Haeger-Aronsen, 1973; Meredith, Moore, and Goldberg, 1974) have emphasized the complex interrelationship that exists between the various metal ions which affect the activity of aminolevulinic acid dehydratase (ALAD; EC 4.2.1.24). Whereas the enzyme is regarded primarily as a useful indicator of lead exposure in man, other metal ions can influence its activity. Abdulla and Haeger-Aronsen (1973) showed that while copper, cadmium, and mercury had a similar effect as lead in inhibiting the human enzyme *in vitro*, the activating effect of zinc (Abdulla and Haeger-Aronsen, 1971) could override this inhibition. Meredith *et al.* (1974) showed a similar effect with rats *in vivo* where zinc and aluminium both had an antagonistic effect on ALAD inhibition by lead.

Previous findings (unpublished) in this laboratory have shown that zinc activation of ALAD varies with pH. The present study is aimed at assessing the effect of zinc on the activity of ALAD inhibited by lead at pH values in the range 6.5-7.5. Lead and zinc were added at concentrations ranging from the normal physiological levels, through those expected in heavy industrial exposure, to non-physiological levels. Results are compared with those obtained by the *in vitro* zinc activation of ALAD in blood from an industrial worker with heavy lead exposure. The significance of the results as they affect both the ALAD assay itself and patients with mixed zinc/lead exposure is discussed.

**Materials and methods**

Heparinized blood was obtained from normal donors with no known exposure either to lead or zinc, and an industrial worker with recent intoxication by lead. Enzyme
activity was determined by the method of Burch and Siegel (1971) modified as follows:

Incubation mixtures comprising:
- 0.5 ml sodium phosphate buffer (0.25M) at pH values ranging from 6.5-7.5 at 0.2 pH intervals,
- 0.9 ml haemolysate (2 ml blood + 13 ml 0.2% Triton X-100),
- 0.2 ml test metal(s) solution or 0.2 ml water (control).

After preincubation for 30 minutes at 37°C, 0.1 ml aminolevulinic acid hydrochloride (0.06m) was added to start the reaction. Incubation was at a temperature of 37°C for 1 hour. The reaction was stopped with 1.3 ml of 10% TCA/0.02m N-ethylmaleimide solution, and porphobilinogen (PBG) was determined in the supernatant at 555 nm using modified Ehrlich's reagent.

Test ions were added to the assay in the form of aqueous solutions of analytical grade lead nitrate and zinc acetate. Lead and zinc were determined in the blood specimens using a Beckman 444 atomic absorption spectrophotometer.

Results

Effect of lead and zinc on ALAD activity in vitro

The mean level of endogenous zinc in all specimens was found to be 0.131 mmol/l, which was within the normally accepted range of 0.11-0.14 mmol/l. The values for endogenous lead fell within the normal range for an urban population (1-0-2.0 μmol/l).

In experiments 1 (A-D) (Fig. 1) exogenous lead and zinc were added at arbitrary concentrations while experiment 1(E) was designed to imitate a hypothetical industrial situation with total lead at a dangerous level and total zinc at double the normal level. This zinc concentration was equal to the highest zinc value obtained by the National Research Institute for Occupational Diseases during a survey of workers at a zinc smelter.

In these five experiments the ALAD activity of the blood specimens used ranged from 287 to 1016 (mean 543) μmol PBG/h per 1 red blood cells (RBC) at pH 6-9, which is of the same order as the series used by Abdulla and Haeger-Aronsen (1973) (526-1166, mean 871). To compare the activating effect of the various zinc concentrations and to eliminate the variation in the control curves over the pH range, all values were expressed (Fig. 1) as a percentage of the corresponding control value at that specific pH.

Lead inhibited ALAD activity in vitro at all the concentrations and over the whole pH range (Fig. 1). Zinc was able to activate the enzyme at all pH values at the various levels tested, the degree of activation being proportional to the exogenous zinc concentration and usually greater at the lower end of the pH range. At the lowest zinc concentration this activation was minimal (Fig. 1C, D). This level of added zinc was equivalent to 0.0054 mmol/l whole blood, however, which was lower than the 0.046 mmol/l which Abdulla and Haeger-Aronsen (1971) claimed to have an inhibitory effect on ALAD.

The antagonistic effect of zinc on the inhibition of ALAD by lead at various pH values.

FIG. 1. The antagonistic effect of zinc on the inhibition of ALAD by lead at various pH values.
In the case of experiment 1(E), zinc was added at an intermediate level equivalent to that encountered in vivo. In that case the antagonistic effect was significant, although the final activity was not equal to that obtained by zinc alone. This experiment showed that a large molar excess of exogenous zinc over lead was not able to antagonize the inhibition of lead completely. This emphasized that the antagonistic effect did not depend on a direct stoichiometric relationship between inhibiting and activating ions.

The in vitro effect of zinc on ALAD activity in a worker with dangerous lead exposure

The previous experiments were extended to investigate the effect of zinc on ALAD estimation in the blood of a worker who had recently been removed from high lead exposure.

A 24-year-old man, showed the following clinical features: blood lead, 4.3 μmol/l blood; PCV, 38; urinary aminolevulinic acid (ALA); 0.27 mmol/l; excessive urinary coproporphyrin; blood ALAD, 131 μmol PBG/h per l red blood cells (at pH 6.9). In this patient endogenous zinc was assumed to be 0.131 mmol/l as results in this laboratory indicated that total blood zinc levels were not affected by changes in lead levels.

ALAD activity was estimated over a pH range of 6.5 to 7.5 with the addition of zinc to the assay mixture at various concentrations as in the previous experiments (Fig. 2). Over the pH range, ALAD was activated in proportion to the amount of zinc added. When, for example, zinc was added to the assay at 0.15 mmol/l whole blood the observed ALAD activity was 3.7 times higher at pH 6.9 than in the control without added zinc. The increased value was equivalent to an ALAD activity of 485 μmol PBG/h per l RBC, which fell well within the range of normal values for this series.

Discussion

The recent observations relating to the antagonistic effect of zinc on lead inhibition of ALAD make it necessary to evaluate the interrelationships in the system in so far as they affect the interpretation of ALAD assay results in a clinical laboratory.

In the present study the antagonistic effect of zinc has been shown to operate dramatically at concentrations in excess of the physiological range. However, this observation is of no clinical significance unless blood samples become contaminated with zinc during collection or before analysis.

When exogenous zinc is added at low concentrations (0.053 mmol/l blood) the antagonistic effect on inhibition of ALAD by lead is minimal, but this addition is equivalent to only a 40% rise from the normal endogenous level of 0.135 mmol/l to a total of 0.188 mmol/l blood.

If, however, zinc is added to the blood in vitro at 0.15 mmol/l blood so as to bring the total zinc in the sample to double the normal endogenous level, a probable physiological situation is obtained which could be expected in vivo under conditions of industrial zinc exposure. In this case we have shown that the ALAD values are appreciably higher than those obtained without the addition of zinc. The degree of antagonism by zinc across the pH range studied here is such that ALAD activity falls within the normal range expected and does not indicate the excessive lead burden that the body is in fact carrying. In this regard also, the pH of assay is important as can be seen in Figs 1(e) and 2, where a shift of pH 0.3 from the standard assay pH of 6.9 can have a marked effect on the degree of antagonism exhibited by the zinc.

The antagonistic effect of zinc on lead inhibition of ALAD activity has been confirmed at concentrations that can be related to the human physiological state. The examination of patients with concomitant lead and zinc exposure may well be complicated by the factors discussed here.

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References


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